

Density Meter Application Note, Concentrations

Density Meter Application Report

Concentration Determinations using the Rudolph Research Analytical DDM 2911 Density Meter

Introduction

From purely an intuitive approach, it is easy to understand how a density meter can determine concentration. If you take a pure water sample, you would find its density to be 0.9982 g/cm³ at 20°C.

Now if you were to take 1g of NaCl (salt) and add this to the same pure water to make a 100 g solution, you would have a density of 1.0053 g/cm³ at 20°C. Add yet another gram of NaCl and the 2% solution now will have a density of 1.0058 g/cm³ at 20°C. If you were to continue this process, you would soon have a complete table of % NaCl vs. density at 20°C. Such a table can be used to determine the concentration of an unknown NaCl solution.

The example above only discusses the concentration determination of NaCl. However, this same type of concentration determination can be done for many other solutions and many such tables already exist in the literature. But it is not too difficult to make new tables for a “two component” or binary solution; where chemical “A” is dissolved into chemical “B”. These can be a solid dissolved into a liquid or a liquid dissolved into a second liquid. An example of the latter is the %Toluene in nHeptane. The DDM 2911 is extremely user friendly.

Rudolph Research Analytical DDM 2911 Density Meter

The DDM 2911 most often provides the fastest, most accurate, and easiest means for determining the concentration of a solution. The common quantitative chemical analysis practice is to titrate a NaOH solution with a standardized HCl solution until a pink end point is reached. But, consider all the various sources of error inherent to this method. The HCl solution is calibrated at 20 °C, but the room temperature is 22 °C, resulting in error #1. Then a standardized HCl solution is used, error #2. The buret is filled to the “zero” line, error #3. Then the titration takes place to the pink end point, error #4. The titrant used, error #5. This titration method is very tedious, labor intensive, and time consuming, and the results are not very accurate. This is why it is common practice for all titrations to be done three times and then the average of those three results is used.

Now consider the numerous advantages to doing this same quantitative analysis by using the DDM 2911 Density Meter. There is no need for the HCl standard solution. Nor is there a need for the phenolphthalein indicator solution. Human error at the pink end point is eliminated. The time required for the complete analysis is only 2 to 3 minutes and the results are very accurate and repeatable. Operator training is minimal as there are only two possible sources of error; contamination of the sample and an incorrect density reading.

bubbles in the sample. Both of these possible errors are completely eliminated because of the exclusive

The DDM 2911 permits the use of a 3 point adjustment calibration. The first calibration point would be c
Subsequently, even the smallest amount of any contamination in the U-tube would be detected by the c
calibration/adjustment points would bracket around the density ranges of the samples to be measured
second source of error, bubbles in the sample, is also completely eliminated by the DDM 2911 exclusiv
bubbles can be easily detected by the operator ensuring precise, accurate measurements every time.

Application Report

Density Meter, Common Applications for the DDM 2911

Density is an important quality control parameter since it is a measure of purity, consistency and also a component or binary solutions.

Concentrations may be determined in the units of %wt, %vol, Proof, Molarity, Normality, ppm, °Brix, API required by the user.

The following industries rely heavily upon these measurements:

Breweries

Solubility of incoming malts, °Balling, °Plato, OG, Specific Gravity, Apparent Extract, consistency of final easily using the Rudolph Research Analytical DDM 2911 Automatic Density Meter.

Chemical Plants

Density is used for a large variety of reasons for both the common bulk chemicals and the specialty and polymers, elastomers, resins, slurries, and more are all measured using Rudolph's DDM 2911 Automat are measured. Dilutions and blending operations are monitored and controlled through the use of density. Density may be used to c Normality, Molarity, percent by weight or volume, PPM, API, and many other related density values.

Distilleries

Density is used as an officially recognized method for the determination of alcohol concentration for qu labeling for declaration of alcohol content for the payment of tax.

Flavor & Fragrances

Density is used to check the blended and/or diluted oils to the desired concentration and customer specifications as alcohol, sugar (°Brix), % solids, and many other parameters. Specific Gravity is always checked on for consistency.

Petroleum

The DDM 2911 complies with all the requirements of ASTM D4052, D5002, D5931, D1250, DIN 51757 Gravity, API Numbers are all determined at various temperatures using the DDM 2911 for QC in Refine

Pharmaceuticals

USP29 <841> Specific Gravity Method II and USP29-NF24 approves the use of the DDM 2911 Automatic Pharmacopoeia 5.0 (2005), pp. 27-28; Pharmeuropa, Vol. 15, 1, January 2003, pp. 174-175 approves Meter. Uses include the checking of incoming raw materials and for final testing of product to be released. The DDM 2911 is compliant with 21CFR Part 1

Soft Drinks – Beverage

The DDM 2911 is used primarily to measure the sugar (°Brix) of the final product to check for proper balance. Consistency of product is essential to this industry and the measurement of density using the DDM 2911 to this end.

Tags: [alcohol](#), [API](#), [ASTM D1250](#), [ASTM D4052](#), [ASTM D5002](#), [ASTM D5931](#), [Automatic](#), [benchtop](#), [Brix](#), [calibration](#), [concentration](#), [DIN 51 757](#), [essential oils](#), [European Pharmacopoeia 5.0](#), [Molarity](#), [Normality](#), [percent by weight or volume](#), [petro chemical](#), [USP29](#), [USP29-NF24](#), [°Balling](#), [°Plato](#)

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